
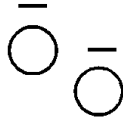


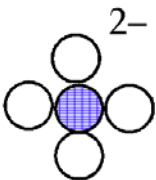
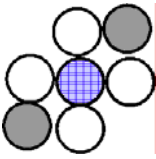


## Unit 4 Worksheet 3

### Representing Ions and Empirical Formulas

	IONS	and	Br <sup>-</sup>	FORMULA	NAME
1.	Ca <sup>2+</sup>  				_____
2.	Fe <sup>2+</sup>		Cl <sup>-</sup>	_____	_____
3.	K <sup>+</sup>  		SO <sub>4</sub> <sup>2-</sup>  		_____
4.	Al <sup>3+</sup>		NO <sub>3</sub> <sup>-</sup>	_____	_____
5.	Pb <sup>2+</sup>		S <sup>2-</sup>	_____	_____
6.	_____		_____	NH <sub>4</sub> OH	_____
7.	_____		_____	KHCO <sub>3</sub>	_____

8. \_\_\_\_\_ and \_\_\_\_\_  $\text{Mg}(\text{NO}_2)_2$  \_\_\_\_\_
9. \_\_\_\_\_ and \_\_\_\_\_  $\text{ZnCO}_3$  \_\_\_\_\_
10. \_\_\_\_\_ and \_\_\_\_\_  $\text{Na}_3\text{PO}_4$  \_\_\_\_\_
11. \_\_\_\_\_ and \_\_\_\_\_ \_\_\_\_\_ silver chromate
12. \_\_\_\_\_ and \_\_\_\_\_ \_\_\_\_\_ lithium chlorate
13. \_\_\_\_\_ and \_\_\_\_\_ \_\_\_\_\_ copper (II) nitrate
14. \_\_\_\_\_ and \_\_\_\_\_ \_\_\_\_\_ iron (III) sulfide
15. \_\_\_\_\_ and \_\_\_\_\_ \_\_\_\_\_ calcium sulfate