

Name _____

Date _____ Pd _____

Chemistry: Unit 4 - Worksheet 4

More Empirical Formulas

IONS	FORMULA	NAME
1. Na^+ and Br^-	_____	_____
2. Cu^+ and SO_4^{2-}	_____	_____
3. Pb^{2+} and Cl^-	_____	_____
4. K^+ and S^{2-}	_____	_____
5. Sn^{2+} and F^-	_____	_____
6. _____ and _____	BaI_2	_____
7. _____ and _____	AlCl_3	_____
8. _____ and _____	$\text{Mg}(\text{NO}_3)_2$	_____
9. _____ and _____	KOH	_____
10. _____ and _____	$(\text{NH}_4)_2\text{SO}_4$	_____
11. _____ and _____	_____	silver oxide
12. _____ and _____	_____	lithium bromide
13. _____ and _____	_____	copper (II) nitrate
14. _____ and _____	_____	magnesium chloride
15. _____ and _____	_____	calcium carbonate
16. Mg^{2+} and NO_3^-	_____	_____
17. Cu^{2+} and OH^-	_____	_____
18. _____ and _____	NaHCO_3	_____
19. _____ and _____	_____	iron (III) sulfide
20. _____ and _____	_____	potassium chromate

Part II

Write the names of the following compounds

1. $\text{Cu}(\text{NO}_3)_2$ _____
2. BaCl_2 _____
3. HgO _____
4. $\text{Ni}(\text{OH})_2$ _____
5. Na_3PO_4 _____
6. CaCO_3 _____
7. CS_2 _____
8. SnBr_4 _____
9. $(\text{NH}_4)_2\text{CrO}_4$ _____
10. $\text{Mg}(\text{NO}_3)_2$ _____
11. Li_2O _____
12. FeS _____
13. NI_3 _____
14. H_2SO_4 _____
15. $\text{K}_2\text{C}_2\text{O}_4$ _____

Part III

Write the formulas for the following compounds

- | | |
|--------------------------------|--|
| 1. copper (II) sulfate _____ | 2. sodium chromate _____ |
| 3. iron (III) chloride _____ | 4. silver sulfide _____ |
| 5. aluminum oxide _____ | 6. zinc nitrate _____ |
| 7. potassium phosphate _____ | 8. strontium fluoride _____ |
| 9. ammonium carbonate _____ | 10. magnesium hydroxide _____ |
| 11. carbon tetrachloride _____ | 12. phosphorus tribromide _____ |
| 13. sulfur hexafluoride _____ | 14. sulfur dioxide _____ |
| 15. chromium (III) oxide _____ | 16. nitric acid _____ |
| 17. hydrochloric acid _____ | 18. lead(II) iodide _____ |
| 19. ammonium nitrite _____ | 20. potassium hydrogen carbonate _____ |