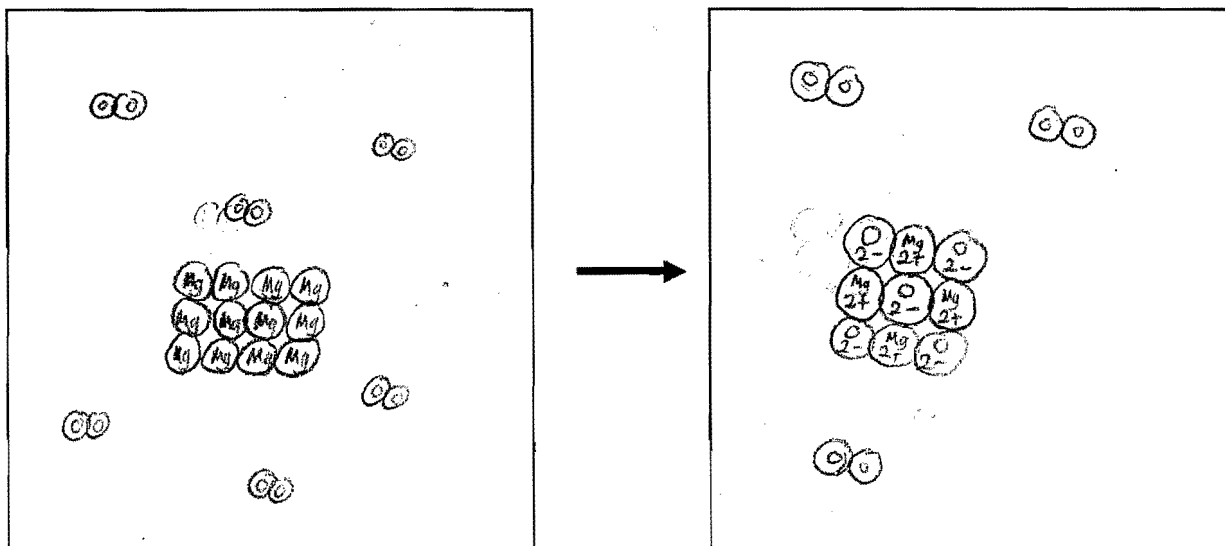


Representing Reactions Quiz

1. As shown in class, a piece of magnesium metal is strongly heated and a white solid is formed. The change is accompanied by the release of energy and a luminous white light.

a) In the spaces below do a before and after drawing at the particle level of the reaction



b) Explain in your own words how the reaction took place and how your representation accounts for the visual observations described above. Make sure to mention the role of each atom participating in the reaction Before the reaction the magnesium and oxygen are neutral.

The reaction took place after the fire sped up the atoms. The magnesium and oxygen started to bump against each other, and the oxygen in the air was able to take electrons from the magnesium. When this occurred the magnesium became positive and the oxygen became negative. The two then attracted each other to form a white solid. The magnesium was oxidated because the oxygen was stronger and able to take the electrons. Any other atoms in the air were spectators.

c) Write the chemical equation which represents the reaction discussed above

